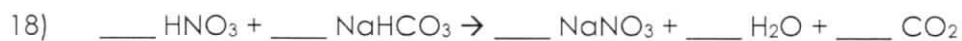
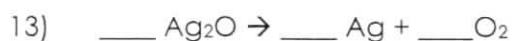
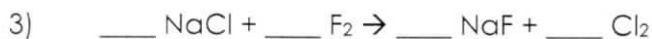


## Chapter 7 Worksheet #1

### Balancing Chemical Equations

Balance the equations below:



## Balancing Reactions Worksheet

Balance the following reactions and identify the type of reaction each represents.

- 1)  $\text{PbO}_2 \rightarrow \text{PbO} + \text{O}_2$
- 2)  $\text{Al} + \text{HCl} \rightarrow \text{AlCl}_3 + \text{H}_2$
- 3)  $\text{Fe}_2(\text{SO}_4)_3 + \text{Ba}(\text{OH})_2 \rightarrow \text{BaSO}_4 + \text{Fe}(\text{OH})_3$
- 4)  $\text{Al} + \text{CuSO}_4 \rightarrow \text{Al}_2(\text{SO}_4)_3 + \text{Cu}$
- 5)  $\text{KClO}_3 \rightarrow \text{KCl} + \text{O}_2$
- 6)  $\text{Mg} + \text{N}_2 \rightarrow \text{Mg}_3\text{N}_2$
- 7)  $\text{FeCl}_2 + \text{Na}_3\text{PO}_4 \rightarrow \text{Fe}_3(\text{PO}_4)_2 + \text{NaCl}$
- 8)  $\text{HgO} \rightarrow \text{Hg} + \text{O}_2$
- 9)  $\text{CaSO}_4 \cdot 2\text{H}_2\text{O} \rightarrow \text{CaSO}_4 + \text{H}_2\text{O}$
- 10)  $\text{Bi}(\text{NO}_3)_3 + \text{NaOH} \rightarrow \text{Bi}(\text{OH})_3 + \text{NaNO}_3$
- 11)  $\text{FeS} + \text{HBr} \rightarrow \text{FeBr}_2 + \text{H}_2\text{S}$
- 12)  $\text{Zn}(\text{OH})_2 + \text{H}_2\text{SO}_4 \rightarrow \text{ZnSO}_4 + \text{H}_2\text{O}$
- 13)  $\text{P}_2\text{O}_5 + \text{H}_2\text{O} \rightarrow \text{H}_3\text{PO}_4$
- 14)  $\text{CaI}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{HI} + \text{CaSO}_4$
- 15)  $\text{Al} + \text{Cl}_2 \rightarrow \text{AlCl}_3$